

Multiple Choice

1. A Chemical Reaction is best described by:
- Production of Heat
 - A white powder going to a clear liquid when heated
 - A Color Change
 - Formation of Bubbles

ANS: B

Multiple Choice

2. A Chemical Equation is best described by:
- Sodium metal reacting with water to form hydrogen gas
 - Description of a chemical change
 - Reactants
 - Products

ANS: B

Multiple Choice

3. Which best describes a chemical reaction:
- There is always a color change
 - Atoms are neither created nor destroyed
 - There is always chemical equation
 - It sometimes involves ionization of salts in water

ANS: B

Multiple Choice

4. An unbalanced equation is best described by:
- The same number of atoms on the right side as the left side
 - Reactants yielding Products
 - $\text{SiO}_2 + \text{HF} \rightarrow \text{SiF}_4 + \text{H}_2\text{O}$
 - The relative numbers of the reactants and products

ANS: C

Multiple Choice

5. The name give to the number placed before each compound formula is:
- The Product Multiplier
 - The Reactant Multiplier
 - Is always an even number

- d. The coefficient

ANS: D

Multiple Choice

6. The book describes a beetle that shoots a boiling water solution. One of the starting materials used to generate the heat is:
- Sodium Hydroxide
 - H_2O_2
 - H_2O
 - Hydrochloric Acid

ANS: B

Multiple Choice

8. What is NOT a driving force for a chemical reaction:
- Formation of a solid
 - Formation of a gas
 - Ionization in an aqueous solution
 - Formation of water

ANS: B

Multiple Choice

9. As described in Chapter 7, which reaction will form a precipitate
- Sodium Metal reacting with water
 - $\text{NaOH} + \text{HCl} \rightarrow$
 - $\text{CH}_4 + \text{O}_2 \rightarrow$
 - $\text{K}_2\text{CrO}_4 + \text{Ba}(\text{NO}_3)_2 \rightarrow$

ANS: D

Multiple Choice

10. An example of a strong electrolyte
- Sugar dissolved in water
 - $\text{CH}_4 + \text{O}_2 \rightarrow$
 - Dissolving table salt in water
 - A gas that has a very bad odor

ANS: C

Multiple Choice

11. The general rule of solubility says
- All salts will dissolve in water
 - When a salt dissolves in water it will ionize
 - Most sulfides will dissolve in water
 - Most nitrates will dissolve in water

ANS: D

Multiple Choice

12. Which chloride will NOT dissolve in water:

- a. Copper Chloride
- b. Zinc Chloride
- c. Silver Chloride
- d. Magnesium Chloride

ANS: C

Multiple Choice

13. Which Hydroxide will NOT dissolve in water:

- a. Sodium Hydroxide
- b. Potassium Hydroxide
- c. Barium Hydroxide
- d. All Alkaline and Alkaline Earth Hydroxides

ANS: C